

Cite this: *J. Mater. Chem. A*, 2020, **8**, 13610

Pressure-driven chemical lock-in structure and optical properties in Sillen compounds PbBiO_2X ($\text{X} = \text{Cl}, \text{Br}, \text{and I}$)†

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Sillen compounds with the general formula of PbBiO_2X ($\text{X} = \text{Cl}, \text{Br}, \text{I}$) are frequently studied as photocatalysts, and have attracted widespread attention due to their degradation of organic contaminants and water oxidation under visible light irradiation. Among many photoelectric materials, the band gap has been reduced approaching to the optimum value by pressure (1.34 eV for photovoltaic materials according to the Shockley–Queisser limit), which is favorable to their photo-responsive applications. However, such enhanced properties are usually restored after the pressure is released. Here, by combining the pressure and chemical engineering tools, we discovered that the optimized structure and optical band gap of PbBiO_2Br could be locked after compression–decompression treatment cycling by selecting appropriate elemental species. The pressure-induced strain retention in PbBiO_2Br could be the cause of the structural and optical irreversible behavior. Moreover, the compression behavior and optical band gap under high pressure in the PbBiO_2X system were studied systematically by *in situ* high pressure synchrotron X-ray diffraction, UV-vis absorption spectroscopy and resistivity experiments. Along with the enhanced photocurrent under compression, the PbBiO_2X compounds under external pressure exhibited great potential for photocatalytic applications under solar light irradiation. In a suitable element category, the pressure-driven structural lock-in preserves its optical performance, opening up a new window for manipulating and filtering better multifunctional materials.

Received 23rd March 2020
Accepted 15th June 2020

DOI: 10.1039/d0ta03291c

rsc.li/materials-a

Introduction

The mixed-anion compound beyond one anionic species in a single phase offers an opportunity to design and control novel physical and chemical properties due to the flexibility of its semiconductor structure and composition.¹ In particular, bismuth-containing semiconductors afford many potential applications in thermoelectric,^{2–5} ferroelectric^{6–9} and photocatalytic fields,^{10–18} owing to the stereoactive $6s^2$ configuration of Bi, as well as the modulated crystal/electronic structure by mixed anions with its distinct electronegativity. For example,

the oxyarsenide BiCuSeO crystallizes in a regular stack of conductive $[\text{Cu}_2\text{Se}_2]$ sheets and insulating $[\text{Bi}_2\text{O}_2]$ sheets. In favor of its layered structure, the intrinsic low thermal conductivity ($0.60 \text{ W m}^{-1} \text{ K}^{-1}$ at room temperature) and the narrow band gap ($\sim 0.8 \text{ eV}$) make it a promising candidate for the highly efficient thermoelectric application.^{4,5} It was found that some Bi-based oxyhalide Aurivillius compounds also adopted an intergrowth of fluorite-like $[\text{Bi}_2\text{O}_2]$ and perovskite-like $[\text{A}_{n-1}\text{B}_n\text{O}_{3n+1}]$ units, such as $\text{Bi}_4\text{Nb}_8\text{O}_8\text{Cl}$, $\text{Bi}_4\text{Ta}_8\text{O}_8\text{Cl}$, and $\text{Bi}_3\text{Pb}_2\text{Nb}_2\text{O}_{11}\text{Cl}$. Their ferroelectricity was widely studied, owing to their polar structure, although the neutron diffraction revealed that a ferroelectric-type phase transition occurred at high temperature from an ideal tetragonal to orthorhombic or pseudo-tetragonal structure.^{8,9} In addition, some other closely related layered oxyhalides, such as the Sillen family, exhibited excellent photocatalytic properties.¹⁸ Under solar light irradiation, they were reported to achieve organic pollutant decomposition,^{14,15} wastewater purification, and water oxidation.^{17,19} So far, the chemical strategies for approaching better performance on the layered Bi-based semiconductors are primarily dedicated to the composition and morphology engineering, all of which are subjected to the fabrication process.

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† Electronic supplementary information (ESI) available: Diffuse reflectance spectra of PbBiO_2X at ambient condition, XRD patterns of PbBiO_2Br and PbBiO_2I at different pressures, GSAS refinement results for PbBiO_2X , bonding distance and angle changes with pressure, pressure dependence of resistivity of PbBiO_2I . See DOI: 10.1039/d0ta03291c

Pressure, as an alternative approach to modify physical/chemical properties in modern material science, has been proven as a clean efficient route to tune the crystal structure and electronic configuration for photoelectric materials.^{20–22} Recently, it has been reported that with an applied external pressure, the bismuth oxyhalide BiOCl suffered from a redistribution of the Bader charge among its component ions. A following isostructural phase transition was experimentally noticed, as proved by the abnormal lattice compression (c/a) within the pressure range from 15.1 GPa to 22.1 GPa.²³ Our previous pressure response studies on bismuth oxychalcogenides have also demonstrated a steady anisotropic lattice contraction upon compression. Due to the pressure-induced competition between the charge delocalization and band gap opening, the electrical resistivity rollback was observed beyond 12 GPa.²⁴ All these findings encourage us to gain deep insight into the structure–properties relationship for other Bi-based mixed anion compounds under high pressure, which has been rarely reported so far.

Sillen X compounds are typical layered semiconductors with single/multiple layers of X atom intercalated in between the fluorite-like Bi₂O₂-based layers.^{25,26} With a single halogen atom layer insertion, PbBiO₂X (X = Cl, Br, I) exhibits potential visible-light-responsive photocatalytic activities in water oxidation or the degradation of organic pollutants due to its appropriate band levels.^{17,27} In this work, by selecting the series X species (X = Cl, Br, I), we conducted the *in situ* comparison studies on the structure and optical property under high pressure with synchrotron X-ray diffraction and optical absorption methods. For the first time, we discovered the band-gap narrowing for all of them during compression, but only PbBiO₂Br retained a partially high pressure structure and property upon decompression. The narrowed band-gap for the pressurized PbBiO₂Br got closer to the pristine PbBiO₂I, which has the narrowest band-gap among these three compounds. From our results, the external pressure could be used as a great potential tool towards better semiconducting materials with engineered band gaps through the irreversible pressure-induced phase transformation. The abnormal structural, optical and transport behaviors of PbBiO₂X could be achieved by combining the chemical and pressure efforts together.

Experimental details

Sample syntheses

As the precursor to synthesize the target PbBiO₂X (X = Cl, Br, and I) samples *via* a two-step procedure, the particulate BiOX reactants were initially prepared by a co-precipitation method as follows: Bi(NO₃)₃·5H₂O (Aladdin, 99.0%) was added slowly into a solution containing stoichiometric amounts of NaCl (Rhawn, 99.5%), NaBr (Rhawn, 99.5%), and KI (Rhawn, 98.5%) respectively, with the Bi : X molar ratio of 1 : 1. The mixture was stirred at room temperature for 0.5 hour in air and washed with distilled water and ethanol, followed by drying at 60 °C for 10 hours. During the process of PbBiO₂X, the well-mixed powders of BiOX and PbO (Aladdin, 99.5%) with the stoichiometric proportion were calcined at 700 °C for 24 hours in alumina crucibles. The phase

identification of the obtained products was carried out on an XPert Powder (analytical) X-ray diffractometer (XRD) with Cu K α radiation.

High-pressure characterizations

Symmetrical diamond anvil cells (DAC) with 300 μm culet-size anvils were employed in all high-pressure experiments, which could facily generate pressure beyond 50 GPa. T301 stainless steel foils were pre-indented to 50 μm in thickness, where the cavities were drilled by infrared laser with 150 μm -diameter. All fine-grinding powders were pressed into pellets with about 10 μm in thickness, and then loaded into the above cavities together with small ruby balls. In this way, ruby was used to calibrate the pressure by the luminescence method. Silicone oil was added in the cavities to serve as the pressure-transmitting medium to provide a quasi-hydrostatic pressure environment. The *in situ* high-pressure angle-dispersive XRD experiments were carried out at the 13 BM-C station of GSECARS at the Advanced Photon Source (APS), in Argonne National Laboratory (ANL). A monochromatic X-ray beam ($\lambda = 0.4340 \text{ \AA}$) was focused to 15 μm in diameter (FWHM) in our diffraction experiments.²⁸ Two-dimensional diffraction patterns were recorded on a charge-coupled device (CCD) detector. The LaB₆ standard powder was used to calibrate the distance between sample and detector and the orientation parameters of the detector. The diffraction patterns were integrated using the program Dioptas.²⁹ It is worth noting that the silicone oil can provide a reasonable quasi-hydrostatic pressure environment below 15 GPa, which may introduce the pressure gradient inside the sample chamber, especially at the high-pressure region. Typically, the pressure variation was less than 0.2 GPa, and no more than 0.5 GPa above 15 GPa. The crystal structures were refined with the Rietveld method by using the General Structure Analysis System (GSAS) program packages.³⁰ *In situ* high-pressure optical measurements were performed on an UV-vis absorption spectrophotometer with the acquisition time of 2 s at room temperature (RT). Each reference was collected on the sample-free area with the same grating size and exposure time. The optical band gap was determined from the Tauc plot, a linear dependence for $(\alpha d h\nu)^2$ versus energy $h\nu$, where α is the absorption coefficient, d is the sample thickness, and h and ν are Planck's constant and frequency, respectively. The resistance measurements under high pressure were carried out by using the four-probe method, which was described in detail previously.²⁴ The resistivity was calculated by the Van der Pauw equation.³¹ For the photocurrent measurements, a Zahner IM6 electrochemical workstation was utilized to record the I - t data by applying a constant bias voltage of 1 V. A 500 W Xenon lamp was utilized as the irradiation source ($\sim 5 \text{ mW cm}^{-2}$ on the sample). The photo response measurements were monitored in the dark and under illumination in the DAC devices.

Results and discussion

Syntheses and crystal structures of PbBiO₂X

As typical Sillen X compounds, they were composed of a single halide anion [X] slab sandwiched by two-dimensional (2D) edge-

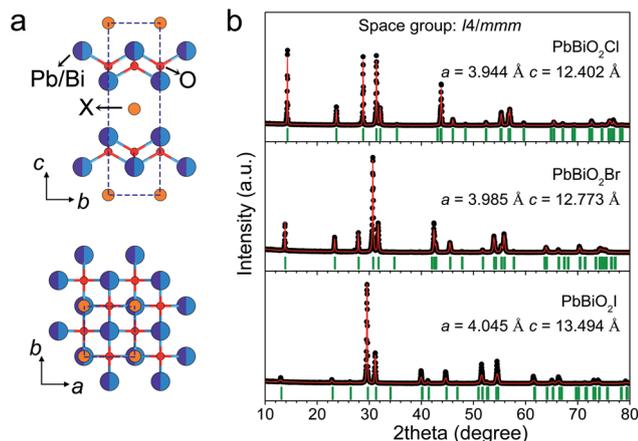


Fig. 1 Crystal structures of PbBiO₂X (X = Cl, Br, I). (a) The tetragonal structure viewed along [100] and [001]. (b) Rietveld refinement results of the PbBiO₂X samples with space group *I4/mmm*. In (b), experimental: black circle; simulation: red lines; Bragg reflections: olive bars.

shared O(Bi/M)₄ tetrahedra units.²⁶ The introduction of the aliovalent cations offers even broader opportunities for the original structural topologies based on the OBi₄ units. CaBiO₂Cl, SrBiO₂X and BaBiO₂X adopted monoclinic (*P2₁/m*) and orthorhombic (*Cmcm*) unit cells, where Bi was partially replaced by an alkaline earth metal.¹¹ Here, Bi is partially replaced by Pb with the same stereoactive 6s² lone pair configuration.²⁵ The natural mineral perite PbBiO₂Cl was reported to crystallize in the orthorhombic structure. However, our XRD examinations showed three profiles of PbBiO₂X matching well with the standard powder diffraction files (PDF #39-0802, PDF #38-1008, PDF #78-0521) with the same tetragonal symmetry *I4/mmm*, as shown in Fig. 1. As the ionic size increased from Cl to I, both lattice *a* and *c* were expanded and the *c/a* ratio increased, since the X ion layers were intercalated between the (Bi/Pb)₂O₂ layers without changing these layers. The refined lattice parameters are: *a* = 3.944 Å, *c* = 12.402 Å for chloride, *a* = 3.985 Å, *c* = 12.773 Å for bromide, and *a* = 4.045 Å, *c* = 13.494 Å for iodide. Tiny deviations of less than 0.2% agreed well with the report by Ketterer *et al.* on them.²⁵

Reversible/irreversible optical property evolution

Hydrostatic pressure can greatly affect the physical/chemical properties of functional materials, including the band gap adjustment. In order to track the optical behavior response to the external pressure, we performed the UV-vis absorption measurements on PbBiO₂X at various pressures. At ambient pressure, PbBiO₂X displayed steep absorption edges at about 517 nm, 538 nm and 563 nm (Fig. S1†) from X = Cl, Br and I, respectively, which are consistent with previous studies.^{12,14,17,27} As the pressure increased, we noticed a piezochromic transformation for chloride from a translucent yellow to dark red, and eventually opaque red at 43 GPa (Fig. 2a). Accordingly, the absorption edges of all samples maintained were gradually redshifted throughout the compression process. The slope of absorbance at the edge tends to be gentle, which could be

associated with the evolution of the band structure, likely, a transformation from the direct to indirect band gap. Meanwhile, over the entire visible range, the absorption increased with increasing pressure. The Tauc plot analysis was achieved by extrapolating the linear portion of the $(\alpha h\nu)^2$ versus energy $h\nu$ plot, indicating a direct band gap of 2.76 eV, 2.64 eV and 2.34 eV from X = Cl, Br, I at the initial pressure point, as shown in Fig. 2c–e. Upon compression, the band gap suffered continuous reduction to 1.98 eV, 1.70 eV and 1.40 eV for PbBiO₂Cl, PbBiO₂Br, and PbBiO₂I, respectively, at the maximum pressure of about 43 GPa. The narrowed band gaps can be explained by the pressure-enhanced orbital coupling of Pb/Bi-6s and X-*np*. Compared with the band gap tuning range of 0.42 eV (from 2.76 eV to 2.34 eV) with the chemical replacement of X from Cl to I, the pressure effect is more significant to ~1 eV (PbBiO₂Cl: 2.76–1.98 eV, PbBiO₂Br: 2.64–1.70 eV, PbBiO₂I: 2.34–1.40 eV) over 43 GPa for each X species. This finding was also realized in other layered bismuth oxychalcogenides.^{24,32,33} Upon decompression, the color of PbBiO₂Cl returned to the original translucent yellow, and the absorption edge completely recovered. The more exotic fact is, after releasing the pressure, unlike chloride and iodide, bromide exhibited a partially retainable band gap of ~2.52 eV, which is ~4.5% less than that of the starting PbBiO₂Br (Fig. 2f and g). The narrow band gap helps to absorb the solar energy more widely. As vital photo-response parameters, both absorption enhancement and band gap narrowing could harvest solar light more effectively.

The optical band gap evolution can be correlated with changes in the structure under pressure. Fig. 3a shows the typical synchrotron XRD patterns of PbBiO₂Cl at selected pressures. All sharp XRD peaks of the samples can be well indexed into the space group *I4/mmm*. As the pressure increases, all the reflections shift to higher 2θ angles, indicating lattice contraction with pressure. No noticeable extra peaks suggest that the initial structure symmetry was retained at the elevated pressure up to 40.5 GPa, the highest pressure applied. A similar structural evolution can also be found for PbBiO₂Br and PbBiO₂I in Fig. S2.† Generally, the peak widths of XRD at high pressure increased, which were typically caused by non-hydrostatic stress and other factors involving the grain size reduction and the increased local strain.^{34,35} However, the peaks (101) for PbBiO₂Cl centered at around 6.62° at ambient pressure abnormally broadened at high pressure, as shown in Fig. 3b. Two-dimensional images showed that the line width of the (101) diffraction peak dissipated gradually with increasing pressure, rather than splitting (Fig. 3c). We deduced that the symmetry did not break here. By using the GSAS package to perform the Rietveld refinement on PbBiO₂X with the same tetragonal structure under high pressure, good agreement factors could be obtained as shown in Fig. S3.†

The lattice parameters of PbBiO₂X based on the Le Bail refinements method are plotted versus pressure in Fig. 4. The *c* to *a* ratio (*c/a*) first decreased markedly with pressure, and then showed a gentle variation as the pressure further increased. Under pressure, the compression of PbBiO₂X was anisotropic with the soft direction along the *c*-axis. In particular, for PbBiO₂I, it reached a minimum (*c/a* dropped about 9.2% from

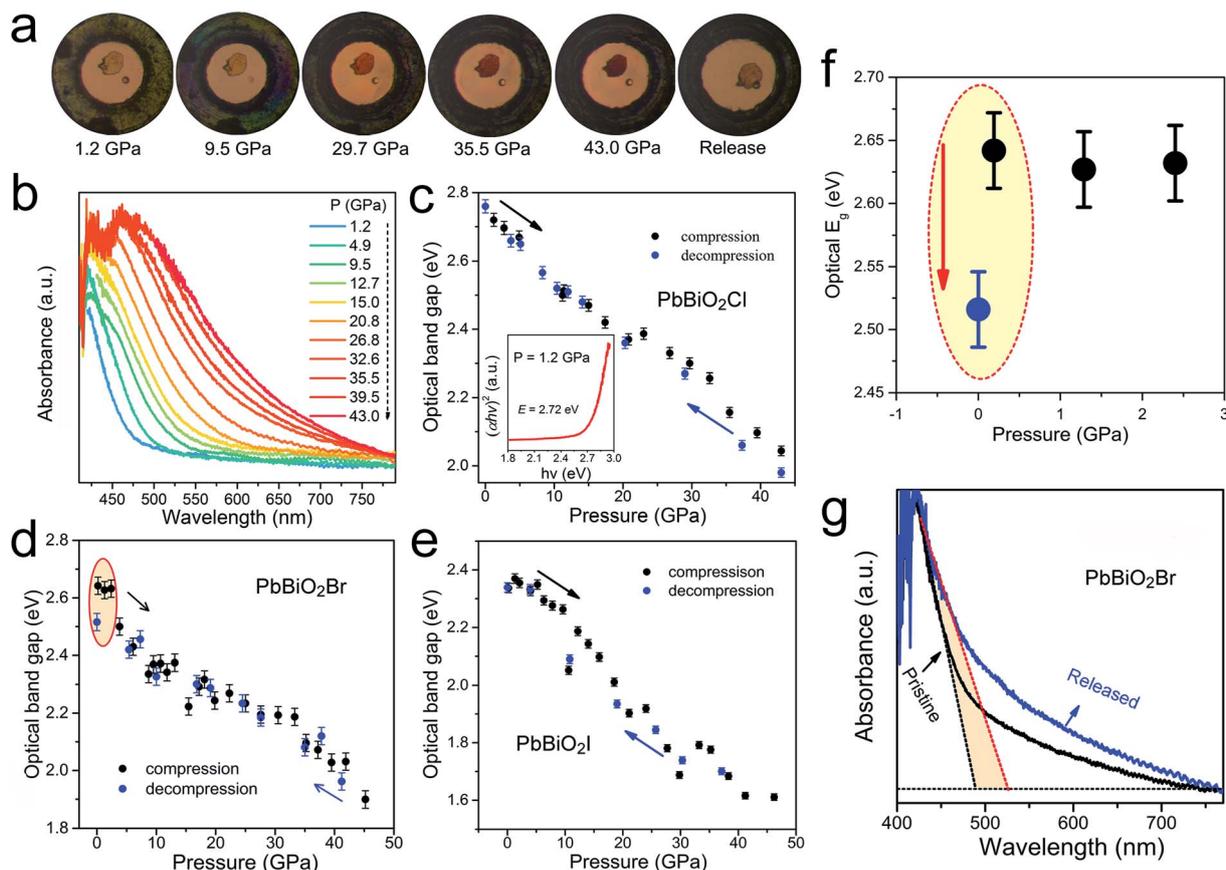


Fig. 2 (a) Optical microphotographs of PbBiO_2Cl inside DAC during compression up to 43.0 GPa and decompression, displaying the piezochromic change from a translucent yellow to dark red. (b and c) Optical absorbance of PbBiO_2Cl upon compression. The absorption edge redshifts with increasing pressure in PbBiO_2Cl . It recovered upon releasing the pressure. The inset gives the direct bandgap Tauc plots of the Kubelka–Munk function at 1.2 GPa for PbBiO_2Cl . The magnitudes of the band gaps can be estimated by extrapolating the linear portion of the Tauc plots to the baselines. (d and e) Pressure dependence of optical band gaps of PbBiO_2Br and PbBiO_2I upon both compression and decompression. (f) The bandwidth before and after compression for PbBiO_2Br . (g) Comparison of the absorbance before and after pressure treatment in PbBiO_2Br .

3.34 to 3.03) at 19.9 GPa. The abnormal relationship of c/a versus pressure for PbBiO_2I could distinctly reflect the isostructural phase transition process, which has been proposed in many similar-layered compounds.^{36,37} The more pressure-sensitive c axis shrinkage results from the considerable interspaces between the $(\text{Bi}/\text{Pb})_2\text{O}_2$ layer and single X atom layer. The rigid c -axis for high pressure phase mainly arises from the stronger interlayer $\text{Pb}(\text{Bi})$ -iodine bonding. Lattice a collapsed slightly at 17.9 GPa for chloride and 10.0 GPa for bromide, as shown in the inset, which we propose is an isostructural change. This mild lattice change could be associated with the evolution of the lone-pair electron configuration of bismuth upon compression. In PbBiO_2X , the $\text{Bi}/\text{Pb}-\text{O}$ and $\text{Bi}/\text{Pb}-\text{Bi}/\text{Pb}$ bonding distances were compressed with increasing pressure as shown in Fig. S4,[†] while the bonding angles for $\text{Bi}/\text{Pb}-\text{O}-\text{Bi}/\text{Pb}$ in the tetrahedron were distorted from $105.60^\circ/117.52^\circ$ to $103.79^\circ/121.54^\circ$, which leads to the compression of the tetrahedral layer. With the non-oxygen anion from Cl to I, the $(\text{Bi}/\text{Pb})_4\text{O}$ tetrahedron is different from the regular tetrahedron, which is attributed to the effect of the halogen atoms with different ionic radii. Next, we fitted the

V - P curves of PbBiO_2X with two sections over the entire pressure range by means of the third-order Birch–Murnaghan equation of state (EoS), as follows $P = \frac{3}{2}B_0 \left[\left(\frac{V_0}{V} \right)^{7/3} - \left(\frac{V_0}{V} \right)^{5/3} \right] \times \left\{ 1 + \frac{3}{4}(B'_0 - 4) \times \left[\left(\frac{V_0}{V} \right)^{3/2} - 1 \right] \right\}$, where V_0 is the initial volume, and B'_0 is the derivative of the bulk modulus with respect to pressure.³⁸ The isothermal bulk moduli B_0 were found to be 56(4) GPa, 53(3) GPa, and 43(3) GPa in the first pressure range for X = Cl, Br, and I, respectively. The reduced bulk module from chloride to iodide corresponds to an enhanced compressibility and decrease of elastic stiffness, as the anions changed from Cl to I. In the higher pressure range, B_0 was fitted as 88(3), 87(2), 98(8) GPa, which are much higher bulk moduli (about 60% higher) than those of the low-pressure phase.

In the Bi-based Sillen compounds, the lone-pair electrons play an important role in their photocatalysis application. The interaction between O-2p and Bi-6s can elevate the O-2p orbitals, favoring the redox reaction in the Bi-based

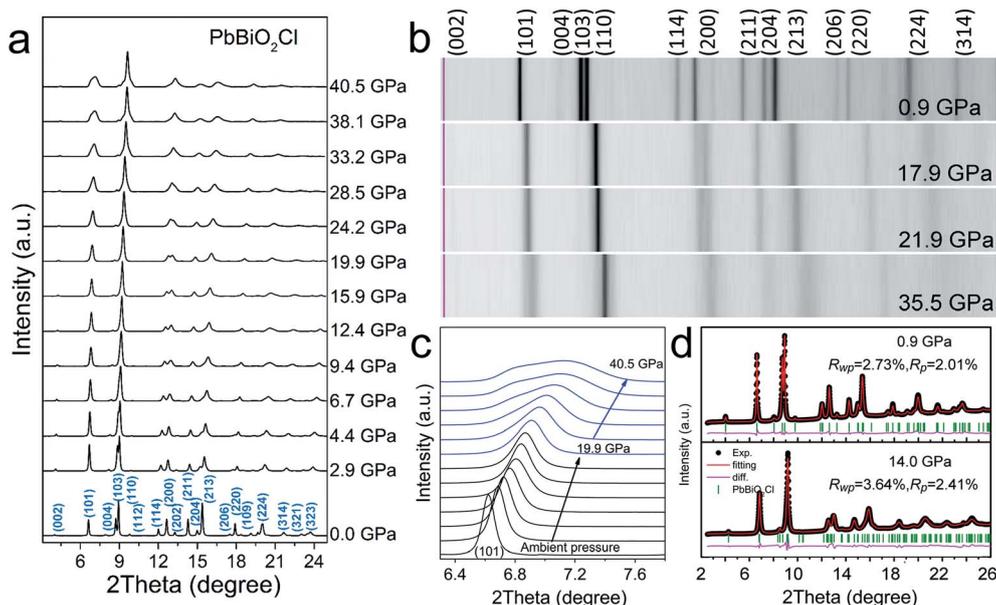


Fig. 3 Selected angle-dispersive XRD patterns of PbBiO_2Cl with incident $\lambda = 0.4340 \text{ \AA}$ at RT (a). (b) 2D images of PbBiO_2Cl at 0.9 GPa, 17.9 GPa, 21.9 GPa and 35.5 GPa, respectively. (c) The evolution of (101) reflections under pressure. (d) Two refinement results for PbBiO_2Cl at 0.9 GPa and 14 GPa.

photocatalyst.³⁹ The phenomenon follows the revised lone pair model proposed by Payne *et al.*⁴⁰ and Walsh *et al.*⁴¹ Most photocatalysis research studies have shown the fact that lone pairs from post-transition metal cations (*e.g.*, Sn^{2+} , Sb^{3+} , Pb^{2+} , and Bi^{3+}) may improve the visible light-responsive photocatalysis efficiency in water splitting applications.¹⁷ Under high pressure, Cheng *et al.* found that perovskite PbCrO_3 underwent an IPT at around 1.6 GPa, accompanying an insulator-metal transition. Based on the XANES measurements, they proposed that Pb-6s, 6p electrons are hybridized with Cr cations at the low-pressure phase. However, after the phase transition, the 6s electrons of Pb were restored to the lone pair state, which caused the insulating-to-metallic transport property change.⁴² Here, we also observed the IPT in PbBiO_2X and presumed it was associated with the pressure-driven lone pair electron configuration evolution. However, there is no conclusive evidence on the effect of the Bi lone pair on the compression behavior. It is indeed intriguing to clarify the role of the lone pair on the structure and properties at high pressure. A continuous development may be required.

Reversible/irreversible structural phase transition

XRD patterns before and after the compression–decompression cycle of PbBiO_2X were then compared, as shown in Fig. 5a–c. The chloride and iodide analogs show good reversibility after pressure releasing. Surprisingly, the crystal structure of the pressure-released PbBiO_2Br exhibits a partial retainable behavior. The XRD peaks from the pressure-released PbBiO_2Br (even after 12 hours relaxation) shifted to a larger 2θ range than those of the pristine sample. In Fig. 5d, the enlarged XRD profiles from the hatched regions in Fig. 5c show clear differences between the pristine and pressure-released PbBiO_2Br

samples. The (002) peak shifted to a higher 2θ angle, along with the (110) and (103) peak merging for the decompressed PbBiO_2Br . By distinguishing them *via* peak fitting, it can be seen that the (103) peak shifts to a higher 2θ angle, while (110) almost completely recovers, meaning that the lattice constant c is largely not recovered due to the irreversible interlayer compression. The unit cell parameters for the pressure-released PbBiO_2Br $a = 3.965(2) \text{ \AA}$, $c = 11.863(3) \text{ \AA}$, and $V = 187.9(3) \text{ \AA}^3$ are smaller than the initial values ($a = 3.988(2) \text{ \AA}$, $c = 12.774(3) \text{ \AA}$, and $V = 203.2(4) \text{ \AA}^3$). This drastically reduced c (7.1%), coupled with an almost totally recovered a -lattice (0.6% contraction), results in a volume reduction of 7.3%. The partially preserved crystal lattice could perfectly explain the abovementioned retainable optical band gap for PbBiO_2Br .

As an alternative approach, pressure can provide energy and tune samples away from its ground state, which allows the system to reach a metastable state with superior properties that could be preserved. Here, we plotted a hypothetical thermodynamic diagram of enthalpy for PbBiO_2X in Fig. 5e. The black and red curves represent the compression and decompression enthalpy change as a function of pressure, respectively. The quenched state (red circle) is more energetically favored when the unit cell is more compressed in comparison to the initial state (black circle). From a thermodynamic point of view, the magnitude of the energy barrier depends on the detailed path of the structure transition. Liu *et al.*⁴³ reported the preserved structures and properties from high pressure treatment in hybrid perovskite $(\text{BA})_2(\text{MA})\text{Pb}_2\text{I}_7$, in which the Pb–I–Pb angle was closer to 180° after the compression and decompression process compared to the as-prepared sample, which enabled a better mixture of the Pb s and I p orbitals. The c -axis in the quenched PbBiO_2Br was

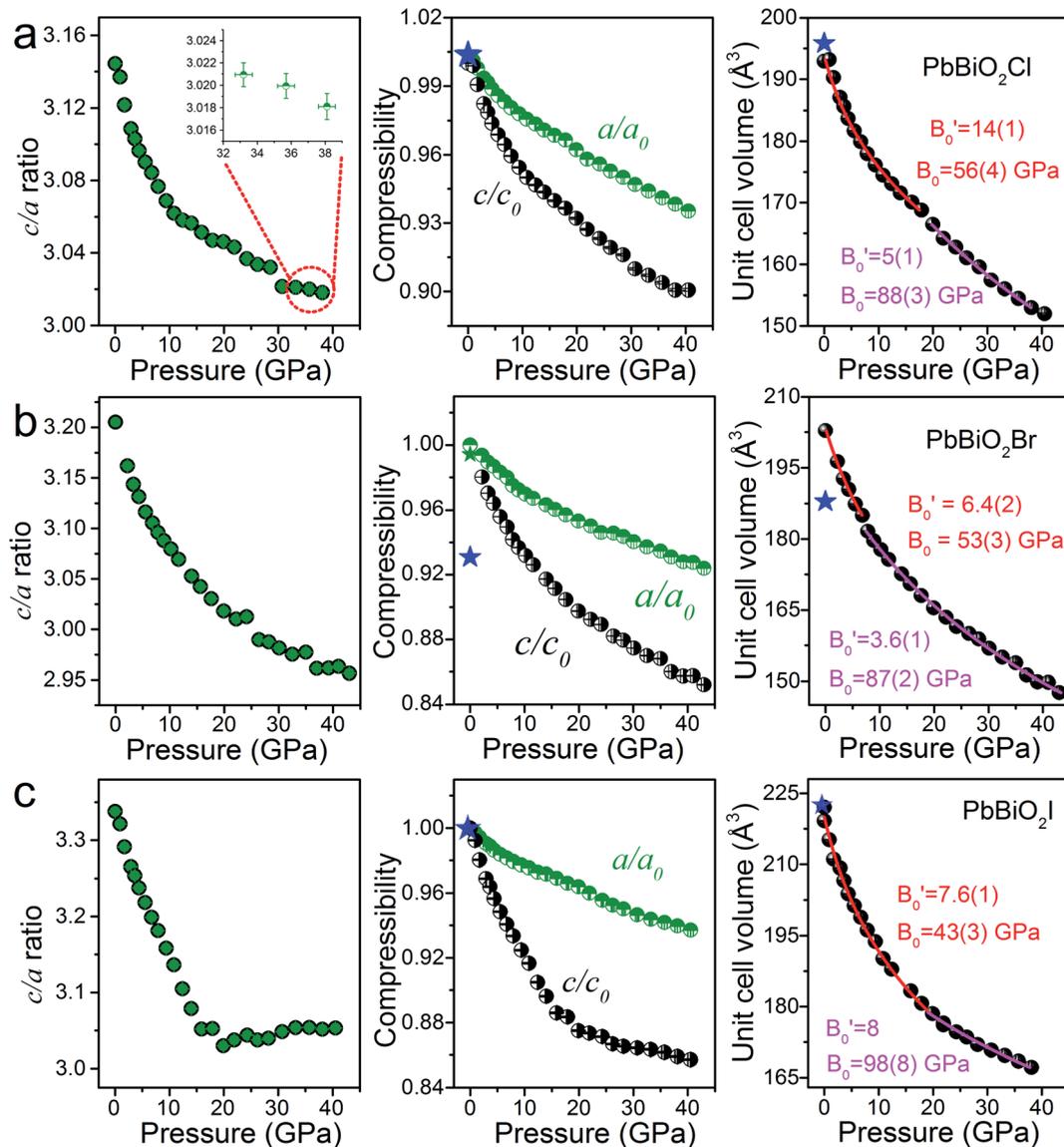


Fig. 4 Evolution of the lattice parameters c/a , the compressibility and unit cell volume of PbBiO₂X as a function of pressure for: (a) PbBiO₂Cl, (b) PbBiO₂Br, and (c) PbBiO₂I. Red and pink lines are the EoS fitting results. Insets give the compressibility c/c_0 and a/a_0 . The error bars, which come from the Rietveld refinement, are covered by the symbols. The enlarged view highlights the size of the c/a error lines for PbBiO₂Cl.

smaller than that in the pristine material, which means the interlayer distance along the c -axis was not totally recovered after being fully released, enabling a strong hybridization of the Pb/Bi 6s–6p and X p orbitals. As a result, a narrower band gap and wider visible light absorbance were achieved. Therefore, we believe that with the proper chemical atoms/ions, pressure can be used as a practical tool to lock-in the structure and physical properties in materials.

Transport properties and photocurrent

In situ high-pressure conductivity measurements were performed on PbBiO₂Br using the quasi-four-probe methods within a DAC to better understand the pressure-tailoring electronic properties. Fig. 6a shows the electrical resistivity evolution of PbBiO₂Br as a function of pressure upon compression

and decompression. The initial high pressure was applied to ensure intimate contact between the powder sample and the Au contact leads. As the pressure increased, the resistivity dropped quickly by up to almost 3 orders of magnitude at 45 GPa. Usually, pressure can enhance the orbital overlap and increase the band dispersion, so the hybridization of the Pb/Bi-6s and O-2p orbitals could cause the monotonic decrease in resistivity.^{24,32} After decompression, the structure changed back to a low-density state, leading to a reversible high resistance. Furthermore, as a reference, we also measured the resistance evolution of PbBiO₂I, which displayed similar behavior to that for the Br-based analog (see Fig. S5†). The conductivities of PbBiO₂X improved significantly under extreme pressure, which may provide some novel application prospects in photo-electrochemical devices.

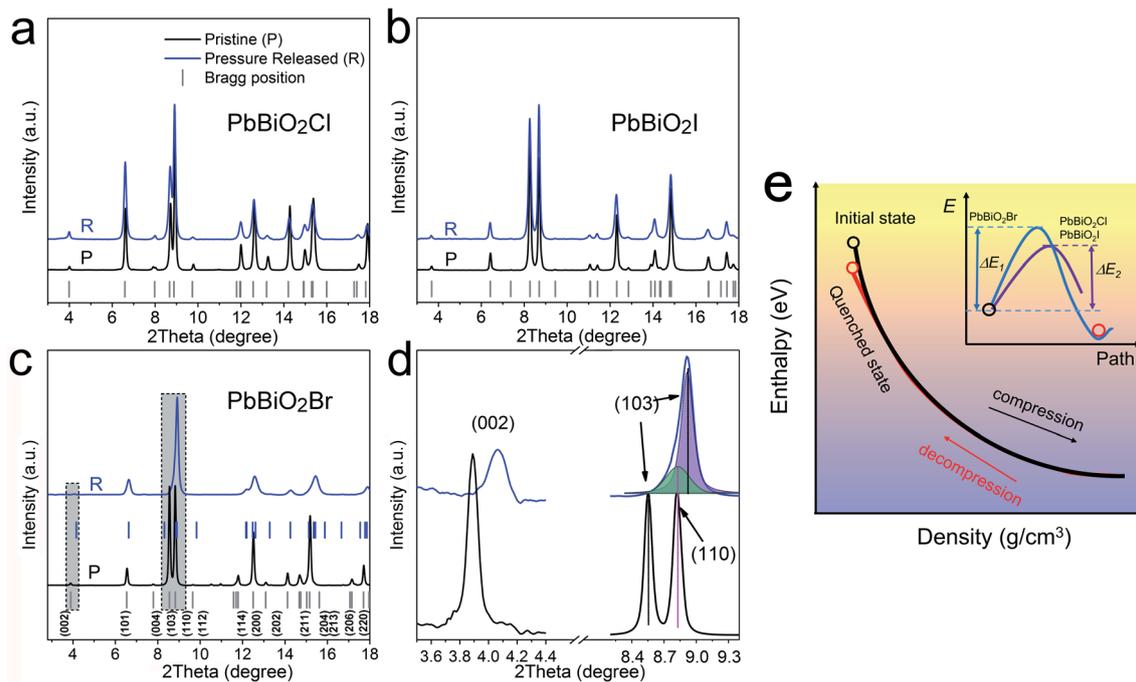


Fig. 5 (a–c) XRD comparison before and after high-pressure treatment up to 43 GPa for PbBiO_2Cl , PbBiO_2Br and PbBiO_2I , respectively. P and R represent the pristine and after-pressure released samples. (d) Enlarged Bragg diffraction peaks (002), (103) and (110) from the hatched regions of (b). The broad peak around 8.9° after pressure releasing was fitted with two individual peaks. (e) The hypothesized enthalpy changes are driven by compression for PbBiO_2X . The reasonably small enthalpy difference ΔE_1 ($\Delta E_1 > \Delta E_2$) between the ground state and quenched state for PbBiO_2Br , leading to its retention structure and properties.

Another essential feature of photocatalysis with photoelectrochemical applications is photoresponsiveness. We evaluated the *in situ* transient photocurrent response for PbBiO_2Br under the simulated solar illumination. As shown in Fig. 6b and c, PbBiO_2Br exhibits a reproducible photocurrent response under high pressure, which follows swiftly with the light on and off, demonstrating the visible light activity and photochemical stability for our sample. The discernable photocurrent under high pressure from 28.7 GPa up to 43 GPa indicated a semiconductor feature as a photovoltaic/photocatalytic material. As the pressure increased, the photocurrent I_{ph} ($I_{\text{ph}} = I_{\text{illuminated}} -$

I_{dark}) was indiscernible under ambient conditions, and was enhanced significantly to $21.5 \mu\text{A}$ ($\sim 48.6 \text{ mA cm}^{-2}$ for $\sim 140 \text{ mW cm}^{-2}$ illumination) at 45.5 GPa. It was more than 4 orders of magnitude higher than the one reported at ambient pressure ($\sim 1.2 \mu\text{A cm}^{-2}$).²⁷ This significantly enhanced photocurrent comes mainly from the remarkable enhancement in conductivity. As discussed above, structural changes under high pressure inherit the layered structural feature, which can explain the persistent photocurrent. As marked with the blue symbols and dashed line in Fig. 6d, the photocurrent reverted with decreasing pressure. In Sillen systems, the pressure-induced

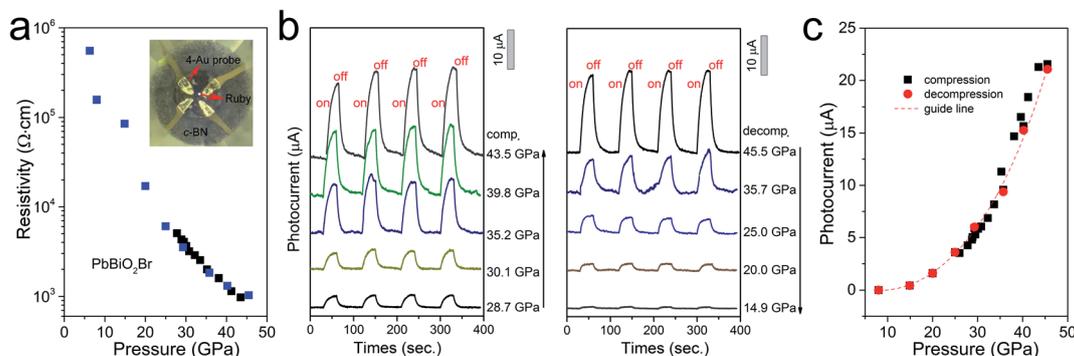


Fig. 6 (a) Resistivity of PbBiO_2Br versus pressure during compression and decompression. Black and blue solid symbols represent the compression and decompression procedures, respectively. (b) Pressure-dependent transient photocurrent responses of PbBiO_2Br under a visible light illumination for several light on–off cycles. (c) Photocurrent I as a function of pressure during compression and decompression.

switchable photoresponsiveness may be closely related to the reversible structure and resistance, and indicates its potential photocatalysis and applications as a switcher or controller.

It is well known that high-pressure research has made unprecedented progress in organic–inorganic halide perovskites by modifying the structure and properties, especially reducing the band gap toward more efficiency in harvesting solar energy for photovoltaic applications.³⁵ In this work, parallel band gap modulation has been successfully spread to complex bismuth-based semiconductors with a layered structure. More importantly, the pressure-induced structural and optical reversibility and irreversibility in Sillen X compound demonstrate a framework for understanding the in-depth structure–property relationships of Sillen compounds, and provide a new dimension for the rational design of semiconductors with tailored applications in optoelectronic/ photoelectrochemical units.

Conclusions

Pressure and chemical engineering tools were applied to the layered semiconductors PbBiO_2X ($\text{X} = \text{Cl}, \text{Br}, \text{I}$) in order to achieve better photoelectrochemical performance. Under compression, three PbBiO_2X complexes underwent a robust tetragonal structural evolution, and presented an anisotropic compression behavior owing to their layered structural feature. PbBiO_2X exhibited continuously narrowed band gaps under compression, which could be associated with the pressure-induced Bi–O bond shortening, and the distortion within the PbBiO_2 layer. Intriguingly, both crystal structures and band gaps of PbBiO_2Cl and PbBiO_2I were recoverable after pressure release. Conversely, PbBiO_2Br showed partial irreversibility in both structural and optical properties. The probable mechanism of the irreversible behavior was also proposed, which was attributed to the unrecovered strains of bromide after pressure-treatment compared to the chloride and iodide analogs. A remarkably narrowed (2.52 eV, 4.5% drop) band gap of PbBiO_2Br compared to the initial value under ambient conditions was obtained after the pressure was removed. By comparison of these three compounds, with the proper atomic size, some structure and optical properties can be locked-in on PbBiO_2X through high-pressure processing. These findings highlight the synergistic effect of pressure and chemical engineering in optimizing the electronic and optical performances, providing new directions for further access to novel multifunctional materials.

Conflicts of interest

The authors declare no competing financial interest.

Acknowledgements

This work was financially supported by the National Nature Science Foundation of China (Grant No. 51527801, 51772184, U1930401) and Science Challenge Project No. TZ2016001. PX^2 program is supported by COMPRES under the NSF Cooperative

Agreement EAR-1661511. APS is supported by DOE-BES, under Contract No. DE-AC02-06CH11357. The authors are also indebted to Dr H. Zhang, J. Liu, F. Zhang, H. Dong for their technical assistance with the UV-vis spectroscopy and XRD experiments.

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